

Arylcyclopropane Photochemistry. A Photochemical Probe into the Stereoelectronic Requirements for Excited State Interaction between Aromatic and Cyclopropane Rings in Arylcyclopropanes [*J. Am. Chem. Soc.*, **97**, 2930 (1975)]. By STEPHEN S. HIXSON* and JOSEPH BOROVSKY, Department of Chemistry, University of Massachusetts, Amherst, Massachusetts 01002.

Arylcyclopropane Photochemistry. The Role of Orbital Overlap Control in the Photochemical Cis-Trans Isomerization of Arylcyclopropanes [*J. Am. Chem. Soc.*, **98**, 2840 (1976)]. By STEPHEN S. HIXSON* and JOSEPH BOROVSKY, Department of Chemistry, University of Massachusetts, Amherst, Massachusetts 01002.

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Peracid Oxidation of Imines. Kinetics and Mechanism of Competitive Formation of Nitrones and Oxaziranes from Cyclic and Acyclic Imines [*J. Am. Chem. Soc.*, **95**, 4692 (1973)]. By YOSHIRO OGATA* and YASUHIKO SAWAKI, Department of Applied Chemistry, Faculty of Engineering, Nagoya University, Chikusa-ku, Nagoya 464, Japan.

In Table I on page 4693, the N's (nitrone) and O's (oxazirane) in line 5 should be interchanged.

Neighboring Group Effect in Heme-Carbon Monoxide Bonding [*J. Am. Chem. Soc.*, **95**, 8475 (1973)]. By C. K. CHANG and T. G. TRAYLOR,* Department of Chemistry, Revelle College, University of California, San Diego, La Jolla, California 92037.

The conclusion in this paper that our model compounds having internally bound pyridine or imidazole do not bind carbon monoxide as well as does myoglobin is incorrect. We have now found that, in contrast to the low binding constant reported in this paper for mesoheme 3-(3-pyridylpropyl)ester mono acid, the carbon monoxide binding constant for the methyl ester form of this compound is about 10^8 M^{-1} (in water), in agreement with the conclusions of M. Rougee and D. Brault, *Biochemistry*, **14**, 4100 (1975).

A Fluxional, Catalytically Active Metal Cluster, Ni₄[CNC(CH₃)₃]₇ [*J. Am. Chem. Soc.*, **97**, 2571 (1975)]. By V. W. DAY, R. O. DAY, Department of Chemistry, University of Nebraska, Lincoln, Nebraska 68508, J. S. KRISTOFF, F. J. HIRSEKORN, and E. L. MUETTERTIES,* Cornell Materials Science Center and Spencer T. Olin Chemistry Laboratory, Cornell University, Ithaca, New York 14853.

On page 2573, footnote 5, line 1, should read: "The low field CH₃ proton NMR peak of relative intensity 27 is at τ 8.16 . . ." Line 4 should read: "9.63 and 8.79, respectively, are also higher than the reported⁶ value of . . ."

Biosynthetic Studies Using ¹³C Enriched Precursors on the 16-Membered Macrolide Antibiotic Leucomycin A₃ [*J. Am. Chem. Soc.*, **97**, 6600 (1975)]. By S. ŌMURA,* A. NAKAGAWA, H. TAKESHIMA, K. ATSUMI, J. MIYAZAWA, Kitasato University and The Kitasato Institute, Minato-ku, Tokyo 108, Japan, F. PIRIOU and G. LUKACS,* Institut de Chimie des Substances Naturelles du CNRS, 91190 Gif-Sur-Yvette, France.

The sentence which begins on the first line on page 6600 should read, "According to the previous paper by Srinivasan and Srinivasan,^{1b} the aglycone carbons of the 16-membered macrolide antibiotic magnamycin B (Ia) are derived, as shown

in Figure 1, from nine acetate, one propionate, and one methionine. In contrast to their conclusion, Grisebach et al.^{1a,c,d} showed that carbons 9 to 16 are derived from acetate, but that other precursors rather than acetate, which are readily formed from glucose, must be responsible for carbons 1 to 6 as well as C-17 and C-18."

The third paragraph on page 6602 should read: "Our results are clearly in contrast to the conclusions of the previous study^{1b} by Srinivasan and Srinivasan in which an acetate origin was assigned to carbons 3 and 4, and are consistent with the finding of Grisebach et al.^{1a} that acetate is not incorporated in this portion of the molecule."

Electronic Relaxation Processes in Retinol and Retinal: Anomalous External Heavy-Atom Effects and Temperature Dependence of Fluorescence [*J. Am. Chem. Soc.*, **98**, 819 (1976)]. By PILL-SOON SONG,* QUAE CHAE, MASAHISA FUJITA, and HIROAKI BABA, Department of Chemistry, Texas Tech University, Lubbock, Texas 79409, and Research Institute of Applied Electricity, Hokkaido University, Sapporo, Japan.

In Table II, column 1, the second compound should read retinal, not retinol.

An Intramolecular Model for the Enzymatic Insertion of Coenzyme B₁₂ into Unactivated Carbon-Hydrogen Bonds [*J. Am. Chem. Soc.*, **98**, 1297 (1976)]. By RONALD BRESLOW* and PYARE L. KHANNA, Department of Chemistry, Columbia University, New York, N.Y. 10027.

Footnote 10 should read:

(10) F. R. Jensen, V. Madan, and D. H. Buchanan, *J. Am. Chem. Soc.*, **93**, 5284 (1971). We have carried out bromination of V at -20, -78, -98, and -160 °C, leading to VI with 48, 28, 17, and 16% ($\pm 5\%$) H on C₁, respectively. Thus $\sim 15\%$ rearrangement occurs on alkylation, and the rest on bromination: *both* steps involve radicals. The other conclusions of the paper are unaffected.

Nitrogen Inversion Rates in Bicyclo[2.2.2]octyl Hydrazines and Amines by ¹³C NMR [*J. Am. Chem. Soc.*, **98**, 1842 (1976)]. By S. F. NELSEN* and G. R. WEISMAN, Department of Chemistry, University of Wisconsin, Madison, Wisconsin 53706.

On page 1850, reference 34 should be:

(34) C. A. Bennett and N. L. Franklin, "Statistical Analysis in Chemistry and The Chemical Industry", Wiley, New York, N.Y., 1954, pp 223-231.

A Synthesis of N-Substituted 1,2-Dihydropyridines [*J. Am. Chem. Soc.*, **98**, 2344 (1976)]. By JOHN N. BONFIGLIO, ILTIFAT HASAN, JOHN J. PIWINSKI, BARRY WEINSTEIN, and FRANK W. FOWLER,* Department of Chemistry, State University of New York, Stony Brook, New York 11794.

On page 2344, line 15, **4c** should read **5c**; line 19, **4d** should read **3**; line 21, **4e** should read "of **4**"; line 22, **4a-c** should read **5a-f**; line 23, **4e** should read "of **4**"; line 24, **4a** and **4d** should read **5a** and **3**; line 33, **4a-c** should read **5a-f**; line 35, **4a-c** should read **5a-f**.

On page 2345, reference 9: **5a-c** should read **5a-f**.

Assessment of the Thioallylic Rearrangement by a Simplified Technique for High-Precision Measurement of Isotope Effects [*J. Am. Chem. Soc.*, **98**, 4009 (1976)]. By H. KWART* and J. STANULONIS, Department of Chemistry, University of Delaware, Newark, Delaware 19711.

Page 4010: Table II, third column, fifth entry should read " 0.9953 ± 0.0027 ".